

LMSSB1

1 2 3 4 5

[Finish attempt ...](#)Time left **0:08:31**

Question 1

Not yet answered

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Study the table below and predict what your unknown would be if it is soluble in all of the three solvents and found to have a density of 0.8 g/mL and its boiling point is near to that of water?

Compound	Density g/mL	B. Point, °C	Water	Cyclohexane	Ethanol
Acetone	0.79	56	s	s	s
2-Butanone	0.81	80	s	s	s
cyclohexane	0.81	83	i	s	s
Ethyl acetate	0.90	77	i	s	s
Heptane	0.68	98	i	s	s
n-Hexane	0.66	68	i	s	s
1-Hexene	0.67	63	i	s	s
isopropanol	0.79	83	s	s	s
Methanol	0.79	65	s	s	s
n-Propanol	0.81	97	s	s	s

Select one:

- A. n-propane
- B. Heptane
- C. 2-butanone
- D. isopropanol
- E. (A) and (B)

[Clear my choice](#)

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Select one:

- A. n-propane
- B. Heptane
- C. 2-butanone
- D. isopropanol
- E. (A) and (B)

[Clear my choice](#)Question **2**

Not yet answered

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What is the possible observation that you may notice upon mixing suitable amounts of  $\text{NaHCO}_3$  (aq.) and  $\text{HCl}$  (aq.)

Select one:

- Precipitate formation
- Appearance of purple color
- Change in odor
- Evolution of a gas.
- Increase in temperature

[Clear my choice](#)[Next page](#)

## مختبر كيمياء عامه وعضوية

## Quiz navigation

[Finish attempt ...](#)Time left **0:05:31**Question **3**

Not yet answered

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Upon mixing of aqueous solutions of  $\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}_{(s)}$  and  $\text{NH}_4\text{NO}_3_{(s)}$ , what would be observation that you may notice?

Select one:

- Decrease in temperature
- Increase in temperature
- Change in color
- Precipitate formation
- Appearance of an odor

[Clear my choice](#)Question **4**

Not yet answered

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One of the following is a signal of having a chemical reaction:

Select one:

- Decrease in solubility of the products in water at lower temperature
- Increase in the boiling point of a liquid substance upon dissolving some salt to it

[Clear my choice](#)

Question **4**

Not yet answered

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Flag question

One of the following is a signal of having a chemical reaction:

Select one:

- Change in the state of a substance from liquid state into gaseous state
- Increase in the volume of a gas upon heating at constant pressure
- None of the answers
- Decrease in solubility of the products in water at lower temperature
- Increase in the boiling point of a liquid substance upon dissolving some salt to it

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## مختبر كيمياء عامه وعضوية

### Quiz navigation



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Time left 0:00:16

### Question 5

Not yet answered

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An unknown compound produces a white precipitate upon mixing with aqueous solution of  $\text{Ba}(\text{NO}_3)_2$ , but produces nothing upon mixing with an aqueous solution of  $\text{BaCl}_2$ . What that unknown would be?

Select one:

- All of the answers
- $\text{ZnSO}_4$
- $\text{Na}_2\text{CO}_3$
- $\text{NH}_4\text{NO}_3$
- $\text{NaCl}$

[Clear my choice](#)

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Quiz navigation



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Time left 0:14:56

Question 1

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Consider the reaction below: ; MM(g/mol.): CH<sub>4</sub>=16, H<sub>2</sub>O=18How many grams of hydrogen gas (H<sub>2</sub>) are produced when 8 g of CH<sub>4</sub> are mixed with 8 g H<sub>2</sub>O?

Select one:

- 2.67 g
- 0.89 g
- 3.0 g
- 1.5 g
- None of the answers

Question 2

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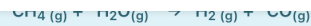
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Time left 0:14:53

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How many grams of hydrogen gas (H<sub>2</sub>) are produced when 8 g of CH<sub>4</sub> are mixed with 8 g H<sub>2</sub>O?

Select one:

- 2.67 g
- 0.89 g
- 3.0 g
- 1.5 g
- None of the answers

Question 2

Not yet answered

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For the reaction between BaCl<sub>2</sub>·2H<sub>2</sub>O and Na<sub>2</sub>SO<sub>4</sub>, to determine the limiting reactant, when the supernatant solution is tested with 0.5 M Na<sub>2</sub>SO<sub>4</sub> solution no precipitate was obtained. Accordingly, which of the following statements is correct?

Select one:

- The limiting reactant is BaCl<sub>2</sub>·2H<sub>2</sub>O
- The limiting reactant is Na<sub>2</sub>SO<sub>4</sub>
- None of the reactants is in excess and the mixing between the reactants is stoichiometric
- None of the answers

## LMSSB1

## Quiz navigation

[Finish attempt ...](#)Time left **0:08:50**Question **3**

Not yet answered

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If equal masses of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  and  $\text{Na}_2\text{SO}_4$  were originally mixed in your unknown mixture, which of the following statements is correct?

MM(g/mol.):  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O} = 244.27$  ;  $\text{Na}_2\text{SO}_4 = 142.05$  ;  $\text{BaSO}_4 = 233.33$

Select one:

- a. The limiting reactant is  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$
- b. The limiting reactant is  $\text{Na}_2\text{SO}_4$
- c. The supernatant solution will give a precipitate upon testing with 0.5 M solution of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$
- d. Both answers (b) and (c) are correct
- e. None of the answers

[Clear my choice](#)Question **4**

Not yet answered

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What is the mass of  $\text{Na}_2\text{SO}_4$  in the mixture if the percentage of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  is 45.0 % and the mass of the mixture used in the experiment is 1.75 g?

Select one:

- None of the answers
- 0.788 g
- 1.09 g
- 0.693 g

[Clear my choice](#)

## مختبر كيمياء عامه وعضوية

### Quiz navigation



Finish attempt ...

Time left 0:08:21

### Question 5

Not yet answered

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Consider the reaction:  $H_2(g) + N_2(g) \rightarrow NH_3(g)$

Mixing of 1.5 moles of each of hydrogen and nitrogen gases gives how many moles of ammonia ( $NH_3$ )?

Select one:

- 1.0 mol.
- 4.0 mol.
- 3.0 mol.
- 2.0 mol.
- 2.7 mol.

[Clear my choice](#)

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LMSSB1

## Quiz navigation

[Finish attempt ...](#)Time left **0:10:36**Question **1**

Not yet answered

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Which of the following statements is true about your recrystallization experiment?

Select one:

- The objective of the experiment was to purify your unknown solid from both soluble and insoluble impurities.
- The objective of the experiment was to separate your unknown solid from a mixture of solids by fractional recrystallization.
- All solid chemical compounds have higher solubility at higher temperatures.
- The method is used for purification of solid chemical compounds from the insoluble impurities only.
- In your recrystallization experiment the solvent is selected such that the substance to be purified has low solubility at high temperature.

[Clear my choice](#)

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the insoluble impurities only.

- In your recrystallization experiment the solvent is selected such that the substance to be purified has low solubility at high temperature.

[Clear my choice](#)Question **2**

Not yet answered

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Which of the following statements is correct about the "Mother Liquor" ?

Select one:

- It is the filtrate obtained after the gravity filtration step, and contains the substance to be purified plus the soluble impurities.
- It is the filtrate obtained after the suction filtration step, and contains the substance to be purified plus the soluble impurities.
- It is the filtrate obtained after the gravity filtration step, and contains the substance to be purified plus the insoluble impurities.
- It is the filtrate obtained after the gravity filtration step, and contains the substance to be purified plus both the soluble and insoluble impurities
- It is the filtrate obtained after the suction filtration step, and contains only the substance to be purified.

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## Quiz navigation

[Finish attempt ...](#)Time left **0:05:31**Question **3**

Not yet answered

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[Flag question](#)

What is the percentage recovery of recrystallization when 2.52 g of pure solid are recovered from a 4.13g sample of your unknown solid?

Select one:

- None of the answers.
- 62.9%
- 78.04%
- 74.11%
- Can't be calculated, more information is needed.

[Clear my choice](#)Question **4**

Not yet answered

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Colored impurities are removed by addition of which of the following?

Select one:

- Charcoal
- Ligroin
- Petroleum ether
- Sodium benzoate
- None of the answers



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Quiz navigation

[Finish attempt ...](#)Time left **0:03:47**Question **5**

Not yet answered

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Which of the following statements is correct about the temperature coefficient of the good solvent for recrystallization?

Select one:

- a. The solvent should be of high temperature coefficient towards the impurities.
- b. The solvent should be of low temperature coefficient towards the impurities
- c. The solvent should be of high temperature coefficient towards the substance to be purified.
- d. The solvent should be of low temperature coefficient towards the substance to be purified.
- e. Both (b) and (c)

[Clear my choice](#)[Previous page](#)[Finish attempt ...](#)