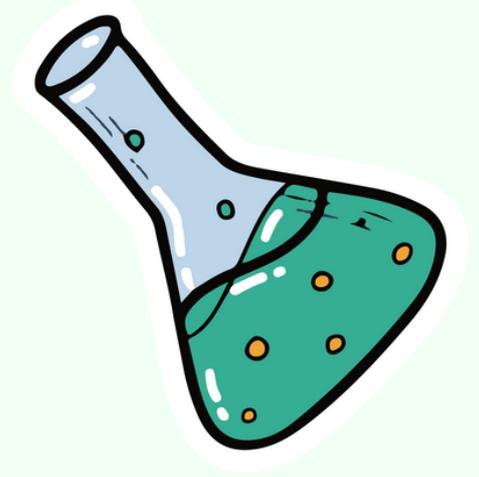


General Chemistry

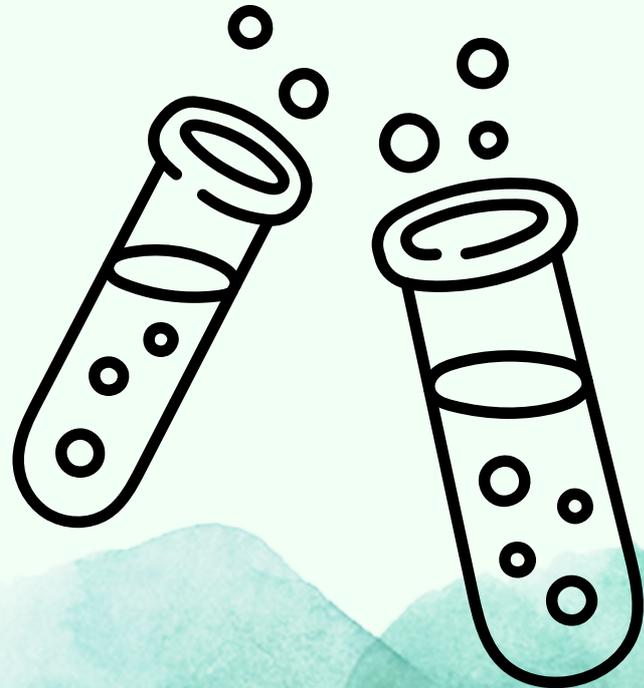
Archive

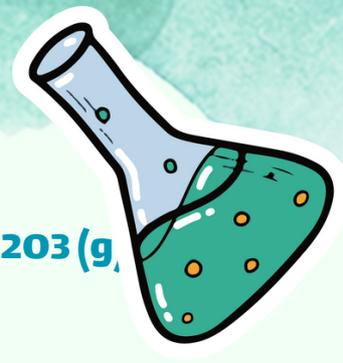
أثر - Mid



Done by: Takwa amjad

Rahma





1-Which of the following is correct about the equilibrium: $3O_2(g) = 2O_3(g)$

- A) $K > K_p$
- B) $Q < K$
- C) $K_p > K$
- D) $K = K_p$

Answer: a

2-What is the $[OH^-]$ for a solution that has a PH of 3.35 ?

- A- $10^{-11} \times 2.24$
- B- $10^{-11} \times 2.01$
- C- $10^{-4} \times 4.47$
- D- $10^{-12} \times 4.4$

Answer: a

3-What is the maximum number of optical isomers for a compound with 4 chiral carbons?

- A)16
- B)12
- C)4
- D)8
- E) None of the answers

Answer: a

4-the number of electrons in the valence shell of aluminum:

- A-1
- B-2
- C-3
- D-4

Answer:c

اللهم قوة، اللهم طاقة، اللهم صبراً





6-which of the following is correct:

A- sulfide (s-2)

B-ammonium chloride (NH₄CL)

C-acetic acid (Hc₂H₃o₂)

D-Barium oxide(Bao)

E-all the following

Answer:e

7-chlorine exists naturally as mixture of chlorine 37 and chlorine 35 an atom of chlorine 37 contains :

A-17p,20n,17 e

B-17p,17 n ,18 e

Answer: a

8-a solutions was labeled 500M NaoH how many grams of NaoH are in 500milli liters of solution:

Answer:10

9-How many milliliter of 1 M Hcl must be added to 50 ml of 500M Hcl to give 0,600M:

Answer:12,5

10-U mix 60ml of 1.0 M silver nitreit with 25.0ml of 0.80M sodium chlorid, what the mass of silver chlorid

Answer: 2.866

اللهم قوة، اللهم طاقة، اللهم صبراً





11-The name of LiNO_3 :

- A-Lithium (I)nitrate**
- B-Lithium nitride**
- C-Lithium nitrite**
- D-Lithium nitrate**

Answer:d

12-Tow isotops of a new discovered element ,the atomic mass of one of them 72.027 g and its aboundancy 36.3 , the average atomic mass is 76.751 , what is the atomic mass to the other isotope ?

Answer:79

13-the oxidizing agent and reducing agent: $\text{SnCl}_2 + \text{FeCl}_3 \longrightarrow \text{SnCl}_4 + \text{FeCl}_2$:

Answer:

SnCl_2 : Reducing agent

FeCl_3 : oxidizing agent

14-A gas mixture of 10g of each of H_2 , N_2 and He under 25 c has a volume of 15,0 L what is the pressure of gas mixture ,and what is the partial pressure of(N_2) gas in mixture:

Answer: 12,82&0,582

15-Tow isotops of a new discovered element ,the atomic mass of one of them 72.027 g and its aboundancy 36.3 , the average atomic mass is 76.751 , what is the atomic mass to the other isotope ?

Answer:79

اللهم قوة، اللهم طاقة، اللهم صبراً





16- $2\text{Cs} + \text{Cl}_2 \rightarrow 2\text{CsCl}$ What happened to Cl_2 ?

- A- oxidation
- B- oxidizing agent
- C- E donate
- D- two of the above

Answer:

17- $27\text{-CuSO}_4 + \text{K}_2\text{CO}_3 \rightarrow \text{CuCO}_3 + \text{K}_2\text{SO}_4$ The pure ionic equation

Answer: $\text{Cu}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CuCO}_3$

18- what is the volume of hydrogen gas produced under STP upon the reaction of 28.3 Mg metal with sufficient amount of hydrochloric acid according to the reaction below



- a) 22.9
- b) 36.3
- c) 67.8
- d) 26.1
- e) 20

Answer: d

19- An unknown gas was found to have a density of 2.053 g/L at STP what is the gas

- H_2
- CH_4
- O_2
- CO_2
- NO_2

Answer: NO_2

اللهم قوة، اللهم طاقة، اللهم صبراً

