Experiment 9:
Reactions of Functional Group - 1
Some Reactions of Hydrocarbons

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Objective:

- To distinguish alkanes, alkenes and aromatic hydrocarbons by their chemical reactions and reactivity.

SEE THE VEDIO:

Reactions Of Aliphatic And Aromatic Compounds,

https://youtu.be/uzME70UbT40

INTRODUCTION

Hydrocarbons

- HCs are organic compounds containing carbon and hydrogen atoms in their structure.

- Hydrocarbons themselves are separated into two types: **aliphatic hydrocarbons and aromatic hydrocarbons.**

a) **Aliphatic hydrocarbons (مركبات خطية مسلسلة)** are hydrocarbons based on chains of C atoms.

- There are **three types of aliphatic hydrocarbons**.
  - ✓ **Alkanes** are aliphatic hydrocarbons with only single covalent bonds between carbon-carbon atoms, they are relatively unreactive , because they have strong c-c single bond.
  - ✓ **Alkenes** are hydrocarbons that contain at least one C–C double bond, more reactive than alkanes
  - ✓ **Alkynes** are hydrocarbons that contain a C–C triple bond, more reactive than alkenes.

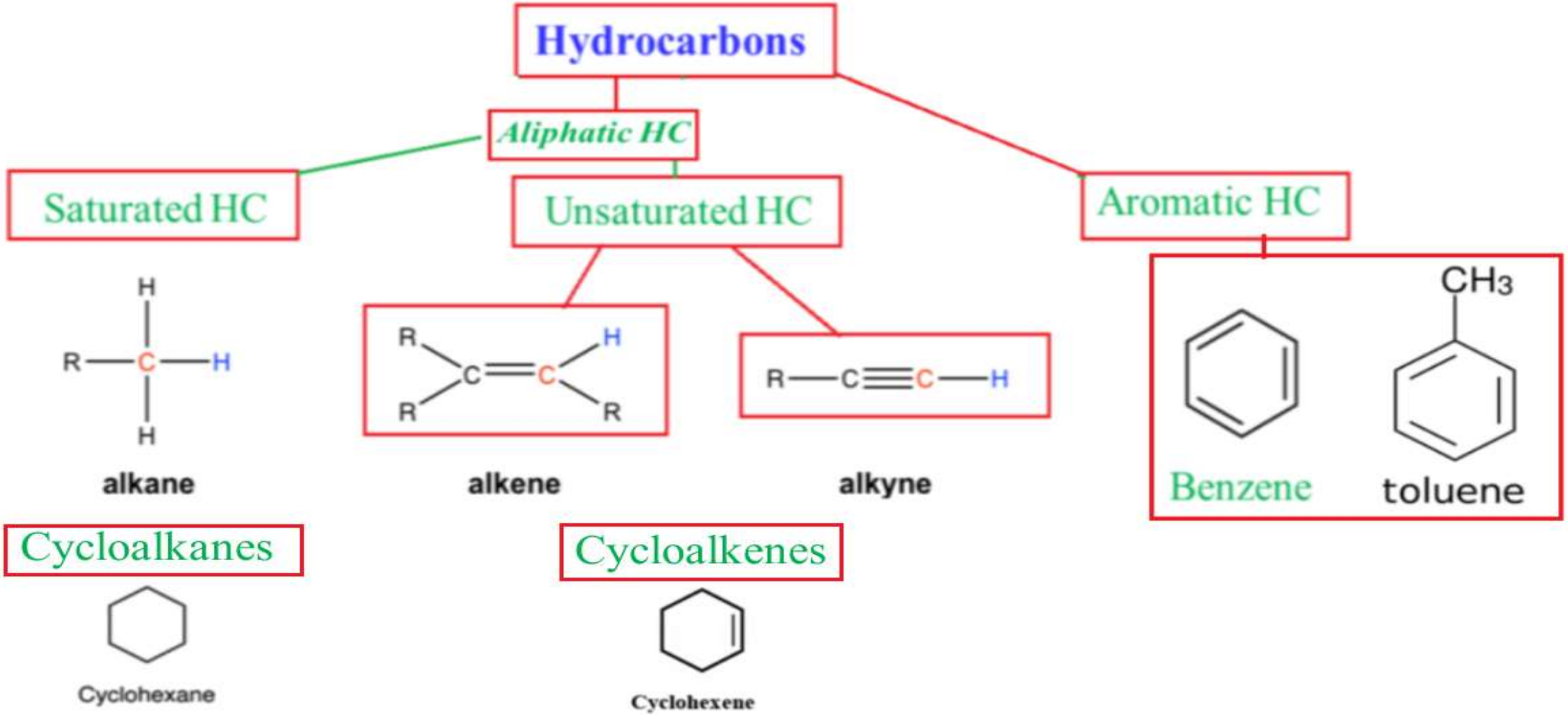
- ✓ **alicyclic hydrocarbons.** Occasionally, we find an aliphatic hydrocarbon with a ring of C atoms; these hydrocarbons are called ***cycloalkanes* (or *cycloalkenes* or *cycloalkynes*).**
- ✓ In general the **reactivity of alicyclic hydrocarbons** are similar to that of linear alkanes and alkenes.

- Because alkanes have the maximum number of H atoms possible according to the rules of covalent bonds, alkanes are referred to as **saturated hydrocarbons**.

b)Aromatic Hydrocarbons

- Aromatic hydrocarbons are sometimes known as "***arenes***" or "***aryl hydrocarbons***".
- Most aromatic hydrocarbons contain a **benzene** ring in their structure; but
- there are non-benzene aromatic hydrocarbons called **heteroarenes**, which follow the "*Huckle's rule*" (Cyclic rings which follow the Huckle's *rule* have 4n+2 number of π-electrons; where n = 0,1,2,3,4,5,6).
- Some aromatic hydrocarbons have more than one ring; they are called **polycyclic aromatic hydrocarbons**.

CLASSIFYING HYDROCARBONS



Reactions of Aliphatic and Aromatic Hydrocarbons

❖ Aliphatic Hydrocarbons:

- **Saturated hydrocarbons** undergo substitution reactions;
- **Unsaturated hydrocarbons** attain the stability by addition reaction. But, some reactions happen under controlled conditions without breaking multiple bonds.

❖ Aromatic Hydrocarbons:

Aromatic hydrocarbons are unsaturated, but have a stable conjugated electron system, so that they are more liable to substitution reactions rather than addition reactions.

Properties of Aliphatic hydrocarbons

- **STRUCTURE AFFECTS MOLECULAR PROPERTIES**

- Alkanes are not polar and are good solvents for other nonpolar molecules
- Alkanes have low reactivity because they are nonpolar and have no charge, and because they have strong single bonds between carbon atoms.
- Alkenes are nonpolar and have low solubility in water. Alkenes are more reactive than alkanes because the double bond increases electron density between the two carbon atoms, providing a good site for chemical reactivity
- Alkynes have physical and chemical properties similar to alkenes but are generally more reactive because the triple bonds cause even larger electron densities than double bonds

Uses of Hydrocarbons

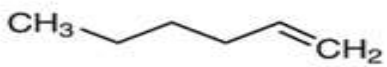
- Hydrocarbons are widely used as fuels. For example LPG (liquefied petroleum gas), CNG (Liquefied natural gas).
- They are used in the manufacturing of polymers such as polyethene, polystyrene etc.
- These organic compounds find their application in the manufacturing of drugs and dyes as a starting material.
- They serve as lubricating oil and grease.

Materials: Test tubes, hexane, cyclohexane, hexene, cyclohexene, m-xylene, dilute $KMnO_4$, aluminum chloride, chloroform, bromine in carbon tetrachloride.

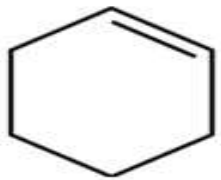
CAUTION!! Be extremely careful handling bromine solution, it can cause burns. If you get it on your skin or clothing wash it immediately with plenty of water.

WEAR SAFETY GLASSES AT ALL TIMES IN LAB

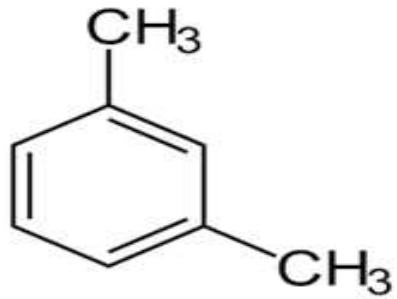
Hexane



1-Hexene

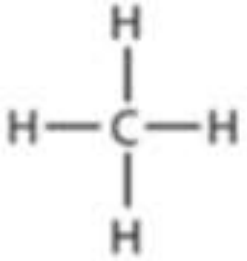


cyclohexene

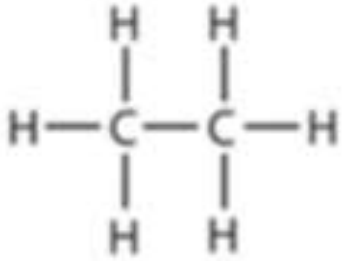


m-Xylene

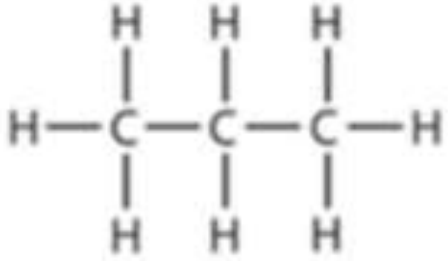
Alkanes and Cycloalkanes



Methane - CH_4

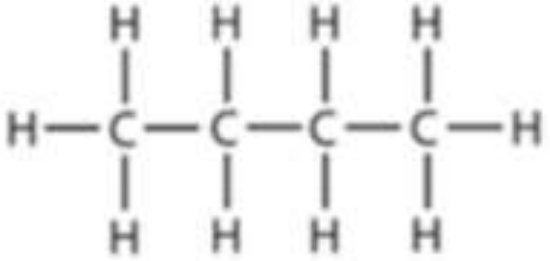


Ethane - C_2H_6

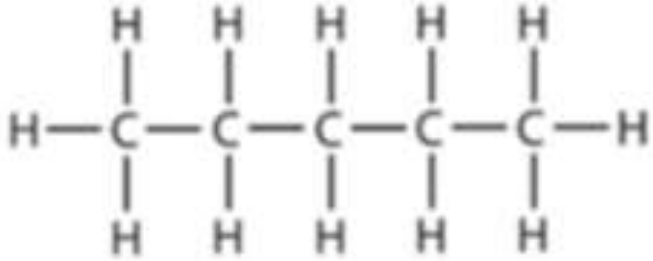


Propane - C_3H_8

Cyclopentane



Butane - C_4H_{10}

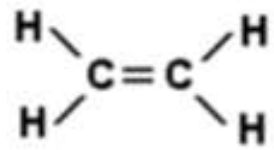


Pentane - C_5H_{12}

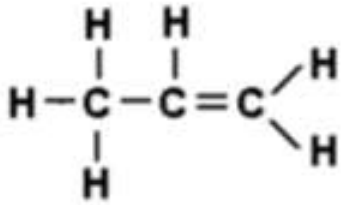
Cyclohexane

Alkenes and Cycloalkenes

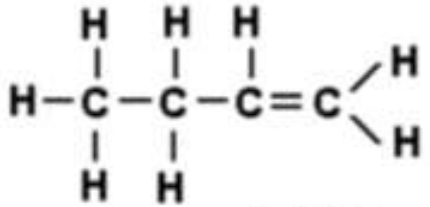
ALKENES



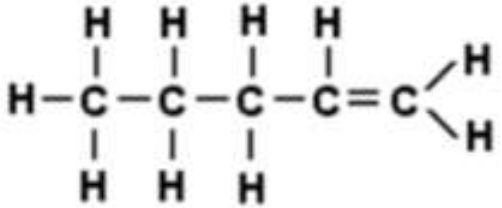
ethene



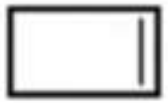
propene



butene



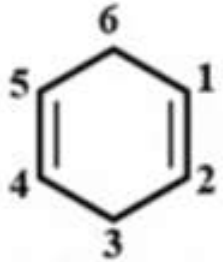
pentene



Cyclobutene

Cyclopentene

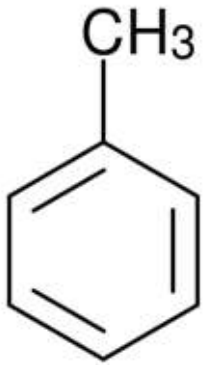
Cyclohexene



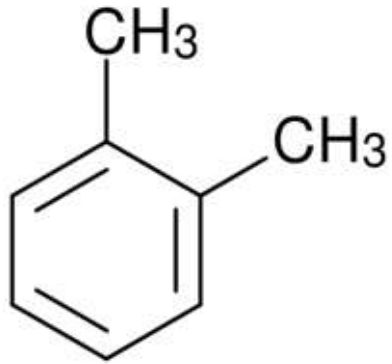
1, 4 - Cyclohexadiene

Aromatic Hydrocarbons

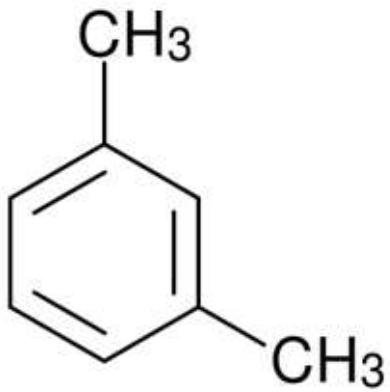
benzene



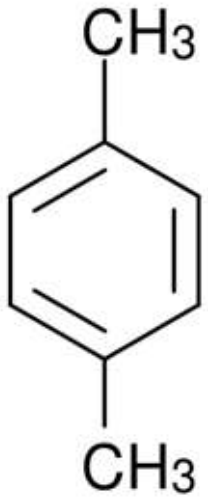
toluene



ortho-xylene



meta-xylene

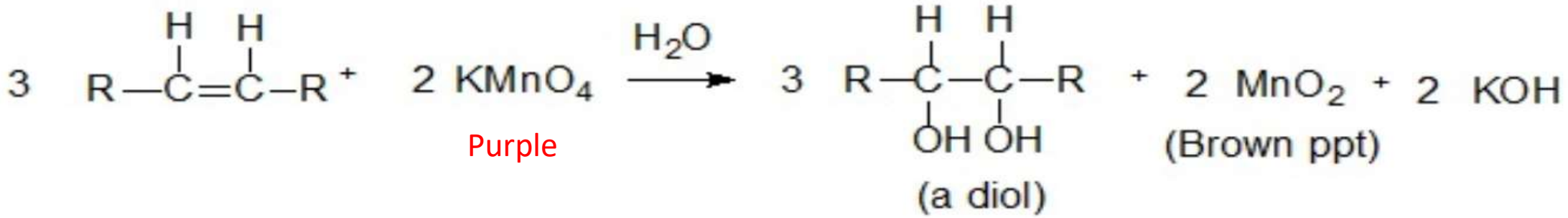


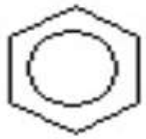
para-xylene

Naphthalene

Part A. Baeyers Test: Reaction of Alkenes with Aqueous Potassium Permanganate

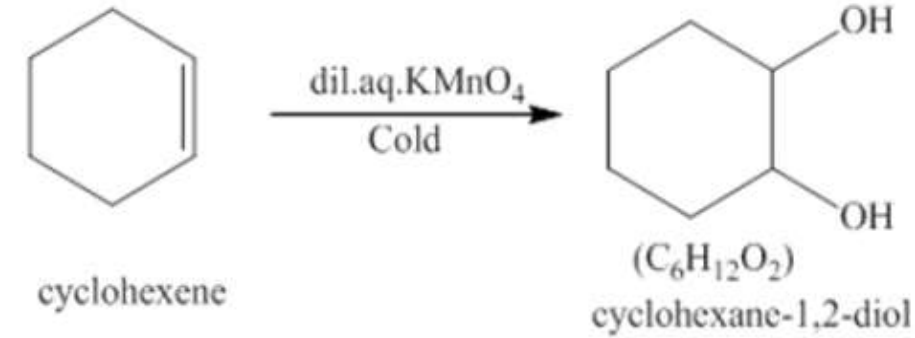
Potassium permanganate is an oxidizing agent that reacts with unsaturated aliphatic hydrocarbons, but does not react with alkanes or aromatic hydrocarbons. The dilute $KMnO_4$ solution has a deep purple color, if there is no reaction you should see no color change. When it reacts with unsaturated aliphatics it produces MnO_2, a brown precipitate. This reaction is useful as a test for the presence of a multiple bond, if there is no other easily oxidizable group, such as an alcohol or aldehyde.

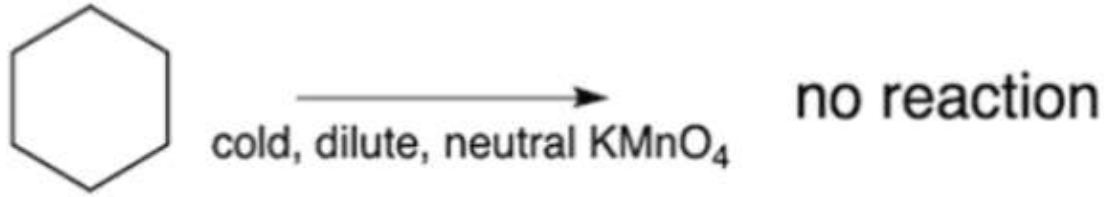


 or $RCH_2—CH_2R + KMnO_4 \longrightarrow$ No Reaction

Procedure:

1. *Place 5 drops of the following HC in clean separate test tubes: Cyclohexane, Cyclohexene, and Toluene.*
2. *Add 2 drops of dilute KMnO4 solution to each test tube and shake.*
3. *Record your observations.*





For Experimental Procedure; See the following YouTubes:

https://youtu.be/B1hqWTKXlQg

$KMnO_4$ test

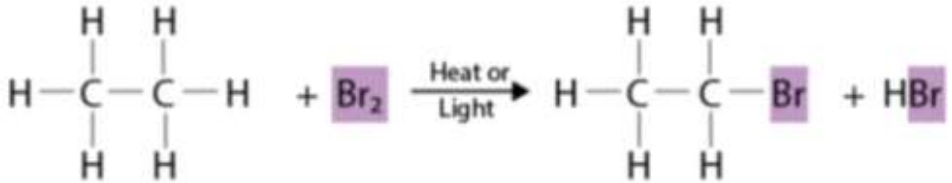
https://youtu.be/pv_-zMbf7Tc

Benzene and Toluene with $KMnO_4$

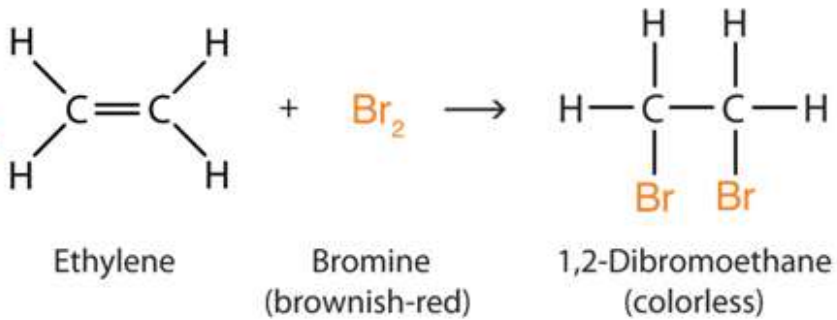
Part B. Reaction of Alkenes with Bromine (Br_2)

Alkenes, but not alkanes or aromatic hydrocarbons will react with Br_2 in solution to produce the corresponding alkyl bromide (or dibromoalkane). The yellow/orange color of Br_2 will disappear as it reacts with the alkenes; the bromoalkane products are usually colorless. Iodine (I_2) gives a similar reaction and is often used to determine the degree of unsaturation of fats and oils. The amount of unsaturation in fats and oils is often given as the iodine number, which is a related to the amount of iodine consumed by a given amount of fat or oil.

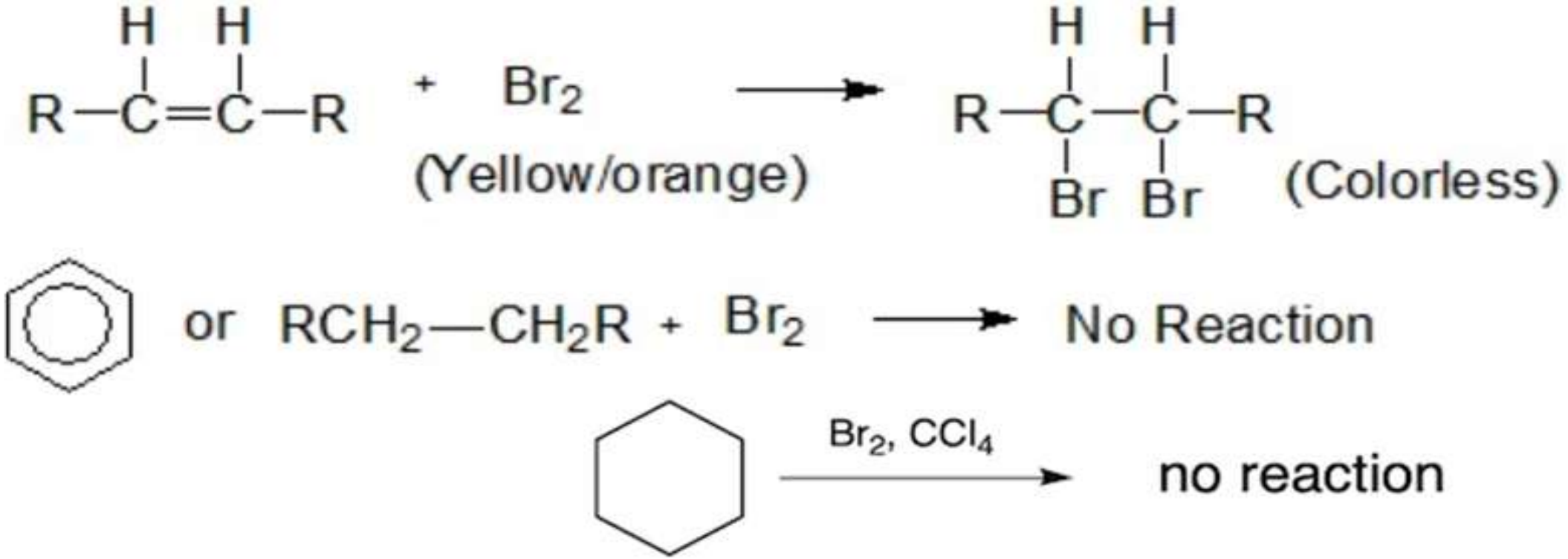
Alkanes react very slowly by substitution only in the presence of ultraviolet light



Alkenes decolorize Bromine rapidly and completely
With dilute solution of Br_2 in CCl_4, even in dark



Ethylene + Br_2 (Bromine, brownish-red) → 1,2-Dibromoethane (colorless)



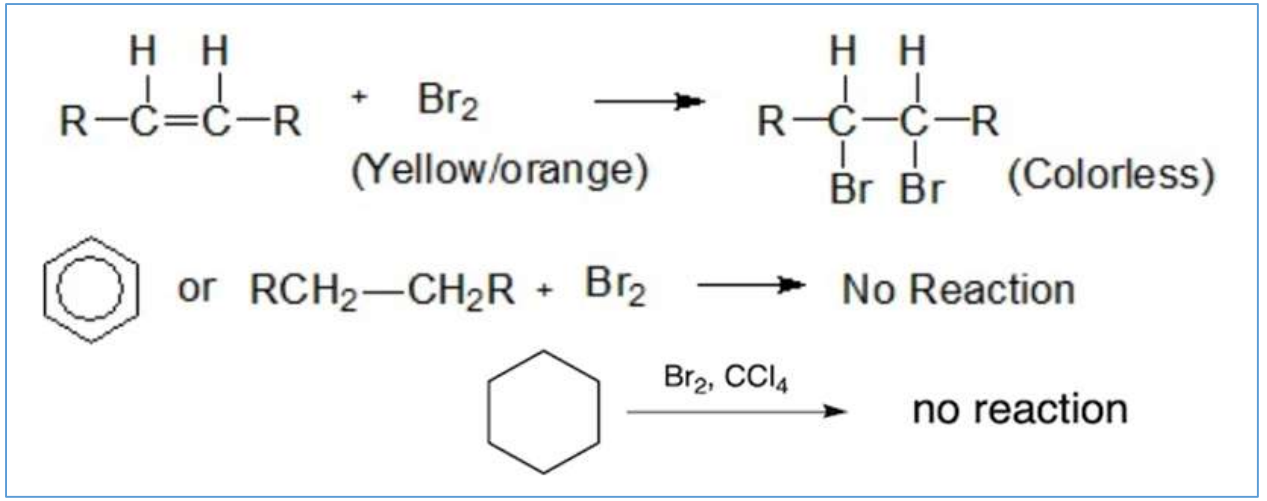
Caution.

-Bromine is highly toxic and can cause severe skin burns. Wear disposable gloves to avoid skin contact with the bromine solution, and do not breathe the vapors. Work in a hood, if possible.

-Sulfuric acid is a corrosive liquid. Avoid skin contact.

Procedure:

1. ***Place 5 drops of the following HC in clean separate test tubes: Hexane, Hexene, and Toluene.***
2. ***Add 2 drops of bromine solution to each test tube and observe the disappearance of red color of Br_2.***
3. ***Record your observations.***



For Experimental Procedure; See the following YouTubes:

https://youtu.be/2C_6ax2TsV8

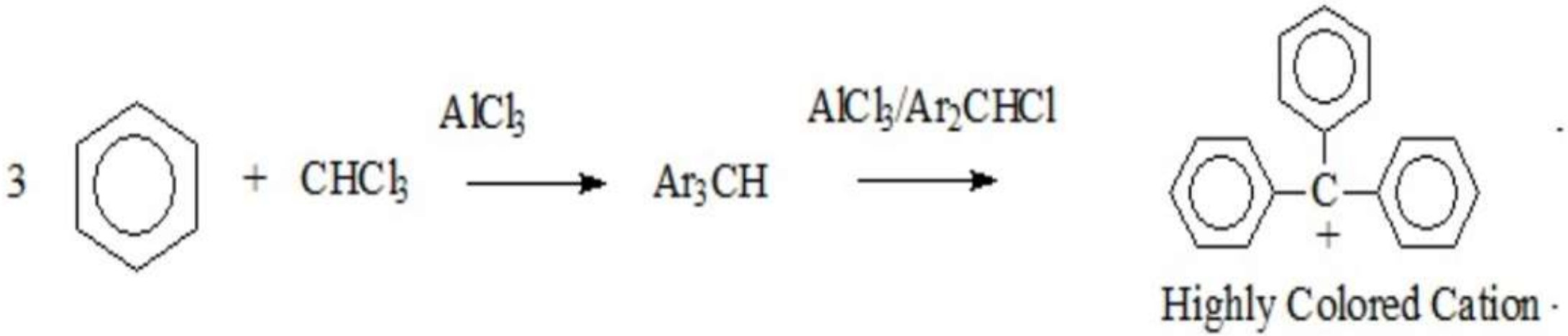
BROMINE TEST HEXANE & HEXENE

https://youtu.be/qEm-CaqhcOs

Br_2 test

Part C. Reaction of Aromatic Hydrocarbons with Chloroform

Aromatic compounds react serially with chloroform in the presence of anhydrous aluminum chloride to produce triarylmethanes (Ar_3CH, where Ar represents an aromatic radical). The product readily undergoes ionization in the presence of $AlCl_3$ and the reaction intermediates to yield a highly colored cation, Ar_3C^+. The color depends on the number of rings in the aromatic hydrocarbon. Benzene and its derivatives give an orange-red color; naphthalene and its derivatives give blue-purple colors.



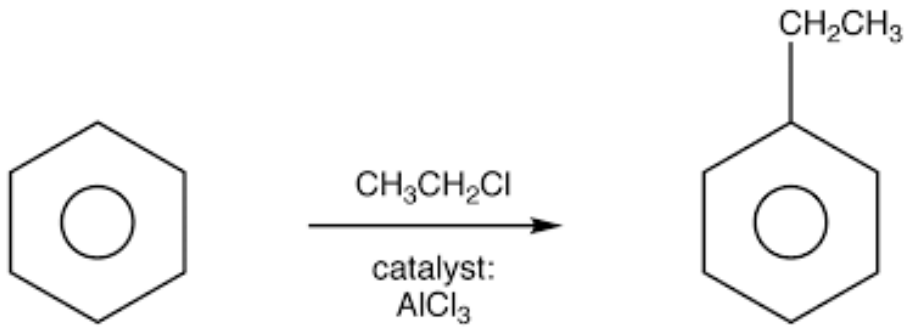
Note: It is essential that the aluminum chloride be anhydrous (water free). Be sure your test tubes and other materials are clean and dry before performing this test.

Procedure

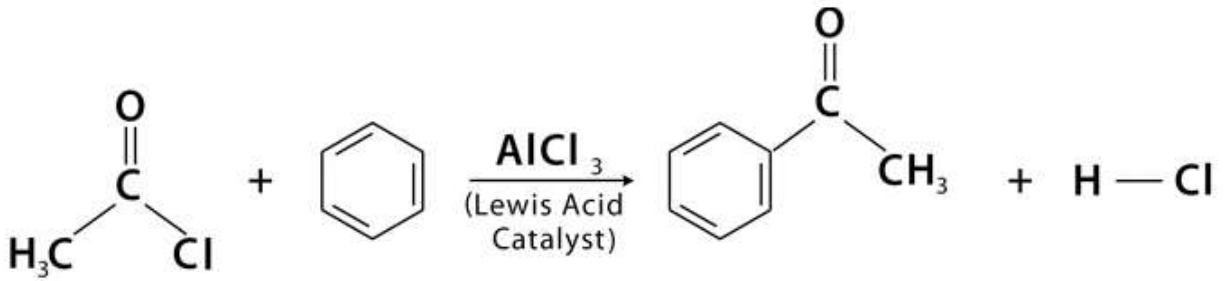
1. Add 1 mL of chloroform to each of 4 (or 5 if you work in pairs) clean, **dry** (no water drops) test tubes.
2. Add 5 drops of the following hydrocarbons to **separate** test tubes containing the chloroform: cyclohexane, cyclohexene, m-xylene, and your unknown. Mix each tube well to dissolve the hydrocarbon in the chloroform.
3. Tilt the first test tube to get some of the solution near the top of the tube, then using a spatula, add a pinch (very little) of $AlCl_3$ to the tilted tube so the powder sticks to the walls of the tube where the solution was (it is not necessary to put the solid $AlCl_3$ in the bottom of the tube).
4. Allow the mixture to stand for 1 or 2 minutes and record your observations on the Report Sheet. Do you observe any color change?
5. Tilt each of the other tubes containing the chloroform mixture with hydrocarbon, and add a pinch of $AlCl_3$ to the tilted test tube so it makes contact with the solution on the walls of the tube as you did for the first tube. Allow it to stand and record your observations on the Report Sheet for each tube.

C. Electrophilic Aromatic Substitution Reactions

1. Friedel – Crafts Alkylation Reaction



2. Friedel – Crafts Acylation Reaction

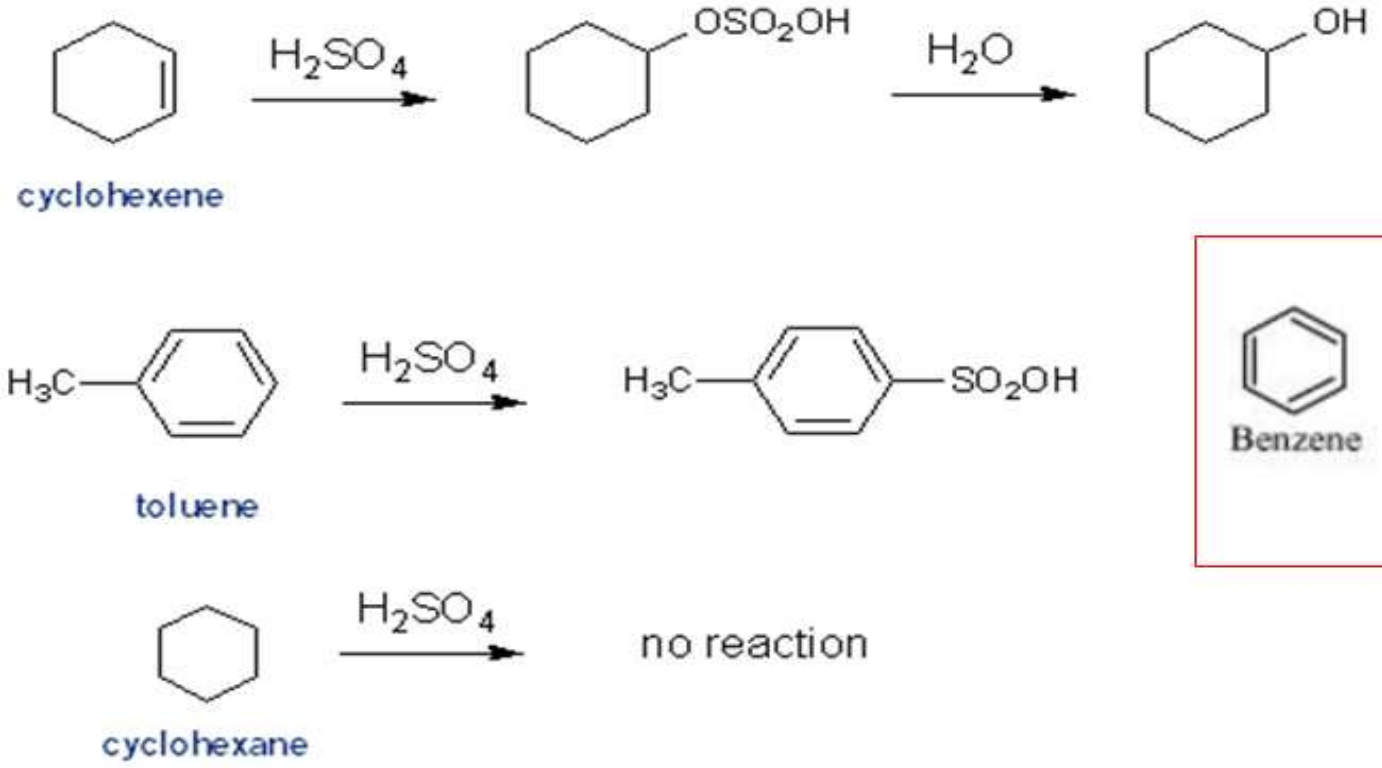


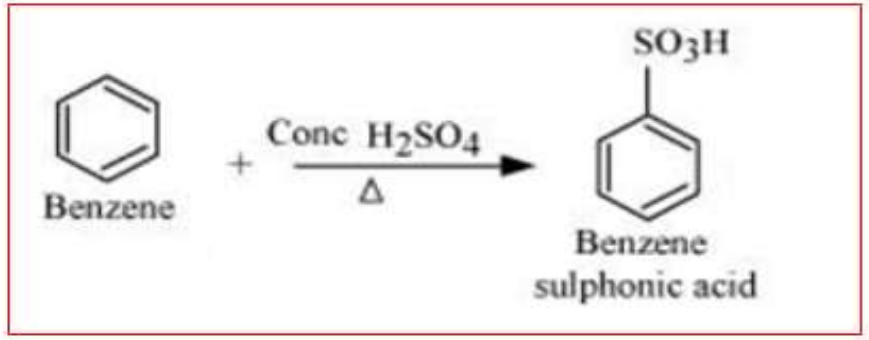
For Experimental Procedure; See the following YouTubes:

https://youtu.be/33D7IYfuEq8
Alkylation and aceylation of benzene

https://youtu.be/mCTHFpwrOKg
Alkylation of *m*-xylene

D. Reaction of Hydrocarbons with Sulfuric Acid





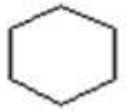
https://youtu.be/oNbwlfqNUD4

H_2SO_4 solubility test cyclohexene, cyclohexane, and benzene

E. Combustion of Hydrocarbon in Presence of O_2

Saturated compounds burn cleanly, while unsaturated ones tend to produce soot.

$C_3H_8 + 5\ O_2 \longrightarrow 3\ CO_2 + 4\ H_2O$

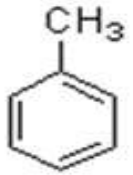


cyclohexane

$C_6H_{12} + 9\ O_2 \longrightarrow 6\ CO_2 + 6\ H_2O + \text{energy}$

cyclohexene

$C_6H_{10} + 8.5\ O_2 \longrightarrow 6\ CO_2 + 5\ H_2O + \text{energy}$



toluene

$C_7H_8 + 9\ O_2 \longrightarrow 7\ CO_2 + 4\ H_2O + \text{energy}$

For Experimental Procedure;
See the following YouTube:

https://youtu.be/EaGbYoZ-6W0

Rxns of Hexane & 1-Hexene:1. Combustion. 2. Br_2 3. $KMnO_4$

Caution.

1. **Sulfuric is extremely corrosive liquids. Avoid contact of these acids with your skin or clothing. Wear disposable gloves. If you have an accidental spill, wash immediately with a large amount of water.**

2. Cyclohexane and cyclohexene are flammable. Do not use an open flame as a heating source during this experiment