

① $v = 1.2 \times 10^{-15} \sqrt{238} = 7.4 \times 10^{-15} \text{ (} v = 1.2 \times 10^5 \text{ (A)}^{1/2} \text{)}$
 Additional: $v = \frac{4}{3} \pi r^2 \rightarrow d = 2v \text{ (radius)}$

② N_B / N_A

$v = 5\% C \rightarrow 0.05 C$

$v_B = 90\% C \rightarrow 0.9 C$

$\frac{v_B}{v_A} = \frac{0.9 C}{0.05 C} = 18 \Rightarrow v_B = 18 v_A$

$\frac{v_B}{v_A} = 19$

③ $\lambda = 9.8 \times 10^{-10} \text{ y}^{-1}$ a) $T_{1/2} = \frac{\ln 2}{\lambda} = \frac{\ln 2}{3.1 \times 10^{-17}} = 2.2 \times 10^8 \text{ s (} 7.07 \times 10^2 \text{ yr)}$

$9.8 \times 10^{-10} \times 1 \text{ yr} \times \frac{365 \text{ day}}{1 \text{ day}} \times \frac{24 \text{ hour}}{1 \text{ hour}} \times \frac{3600 \text{ s}}{1 \text{ second}}$

$\lambda = 3.1 \times 10^{-17} \text{ s}^{-1}$

b) $N_0 \rightarrow 1 \times 10^6 \text{ g}$ (initially at $t=0$)

$N \rightarrow 235 \text{ g}$

$N = 6.022 \times 10^{23}$

$N_0 = \frac{1 \times 10^6 \times 6.022 \times 10^{23}}{235} = 2.5 \times 10^{15}$
 Atomic mass for U is $A = 235$ so, 235 g will contain 6.022×10^{23} nucle. Avogadro number

$R = -\lambda N \Rightarrow N = \frac{R}{\lambda} = \frac{8.1}{3.8 \times 10^{-12}} = 2.12 \times 10^{12}$

$N = N_0 e^{-\lambda t} = \frac{N}{N_0} = e^{-\lambda t}$

$\ln \frac{2.12 \times 10^{12}}{2.5 \times 10^{15}} = -(3.8 \times 10^{-12}) t$

$t = 7.536 \times 10^2 \text{ s (} 238, 960.36 \text{ yr)}$

c) $N = N_0 e^{-\lambda t}$

$= 2.5 \times 10^{12} \times e^{-(3.1 \times 10^{-17})(3.1 \times 10^{16})}$
 $= 9.56 \times 10^{11}$

$\Rightarrow t = 10^4 \text{ yr} \times \frac{365 \text{ day}}{1 \text{ day}} \times \frac{24 \text{ hour}}{1 \text{ hour}} \times \frac{3600 \text{ s}}{1 \text{ second}}$
 $t = 3.1 \times 10^{16} \text{ s}$

④ $T_{1/2} = 10 \text{ min} \times \frac{60 \text{ s}}{1 \text{ min}} \Rightarrow T_{1/2} = 600 \text{ s}$

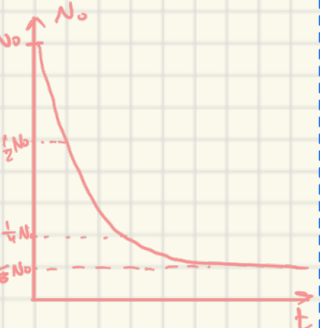
$\lambda = \frac{\ln 2}{T_{1/2}} = 1.155 \times 10^{-3} \text{ s}^{-1}$

Maximum possible number of Cu = N_0

90% of Maximum possible number = $0.9 N_0 = N$

$N = N_0 e^{-\lambda t}$
 $0.9 N_0 = N_0 e^{-\lambda t}$

$\ln 0.9 = -\lambda t \Rightarrow t = \frac{-\ln 0.9}{1.155 \times 10^{-3}} = 91.22 \text{ s} = 1.5 \text{ min}$



⑤ $T_{1/2} = 18.72 \text{ day} \times \frac{24 \text{ hour}}{1 \text{ day}} \times \frac{3600 \text{ s}}{1 \text{ hour}}$

$T_{1/2} = 1617408 \text{ s (Th)}$

$\lambda = \frac{\ln 2}{T_{1/2}} = 4.2 \times 10^{-7}$

$T_{1/2} = 11.42 \text{ day} \times \frac{24 \text{ hour}}{1 \text{ day}} \times \frac{3600 \text{ s}}{1 \text{ hour}}$

$T_{1/2} = 987552 \text{ s (Ra)}$

$\lambda = \frac{\ln 2}{T_{1/2}} = 7.01 \times 10^{-7}$

$N_0(\text{Th}) = 10^6$
 $N_0(\text{Ra}) = 0$ at $t=0$

After $t=15 \text{ d} \rightarrow t = 1296000$

For Th $\rightarrow N = N_0 e^{-\lambda t} = 10^6 e^{-(4.2 \times 10^{-7})(1.296 \times 10^6)} = 0.574 \times 10^6 \text{ Nuclei}$

For Ra $\rightarrow N = N_0 e^{-\lambda t}$

$\frac{N}{N_0} = e^{-\lambda t} = e^{-(7.01 \times 10^{-7})(1.296 \times 10^6)}$

$\frac{N}{N_0} = 0.403 \Rightarrow N = 0.403 N_0$ at $t=15 \text{ day}$

⑥ For $^{14}\text{C} \rightarrow T_{1/2} = 5730 \text{ yr}$ (مغز)

$\lambda = \frac{\ln 2}{T_{1/2}} = \frac{0.693}{5730 \times 3.156 \times 10^7} = 3.8 \times 10^{-12} \text{ s}^{-1}$

$^{14}\text{C} \Rightarrow 14 \text{ g} \rightarrow 6.02 \times 10^{23}$
 $175 \text{ g} \rightarrow N_0$

$N_0 = 75.25 \times 10^{23} \text{ Nuclei}$

$N = N_0 e^{-\lambda t} = \frac{N}{N_0} = e^{-\lambda t}$

$\ln \frac{2.12 \times 10^{12}}{2.5 \times 10^{15}} = -(3.8 \times 10^{-12}) t$

$t = 7.536 \times 10^2 \text{ s (} 238, 960.36 \text{ yr)}$

⑦ 350 rad from α radiation

1 α particle $\rightarrow 20 \text{ x-ray}$

350 $\alpha \rightarrow X$
 $X = 7000 \text{ rad}$

Type	RBE
α and γ rays	1
Protons	1
Slow neutrons	2
Fast neutrons	10
particles and heavy ions	20

The 350 rad of α particle radiation does the same damage as 7000 rad of x-ray

⑧ 4 s.v of γ ray

Effective dose (Sv) = Dose (Gy) \times RBE
 $4 \text{ (Sv)} = \text{Dose (Gy)} \times 1$
 Dose (Gy) = 4

\therefore Number of sievert = Number of gray

⑨ 1 Fast neutron $\rightarrow 10 \text{ x-ray}$ (From the table)

72 rads of fast neutrons $\rightarrow X$

$\approx 72 \text{ rads of fast neutrons cause biological damage equivalent to } 720 \text{ rad of x-ray.}$

1 slow neutrons $\rightarrow 5 \text{ x-ray}$ (From the table)

$X \text{ rads of slow neutrons} \rightarrow 720 \text{ rads x-ray}$
 $\approx 144 \text{ of slow neutrons makes the same damage as } 72 \text{ rads of fast neutrons.}$

⑩ 1 gray = 1 J/kg

4.56 gray = Energy \Rightarrow Energy = 4.5 J/kg
 For a body with mass of 65 kg \Rightarrow Energy = 4.5(65) = 292.5 J